

Synthesis and X-ray structures of cadmium-containing dinuclear double helicates derived from ligands containing *N*-oxide units

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The syntheses of two novel polydentate pyridyl-thiazolyl ligands which contain an *N*-oxide unit on the central pyridyl ring are reported. Reaction of either of these ligands with $\text{Cd}(\text{ClO}_4)_2 \cdot 6\text{H}_2\text{O}$ results in the formation of a dinuclear double helicate. The solid state structures of these transition metal helicates demonstrate the ability of the *N*-oxide unit to partition the ligands into two separate binding domains.

1. Introduction

Formation of transition metal helicates from the self-assembly of organic ligands and transition metal ions has received considerable attention in recent years.^{1–6} One of the main considerations for the successful self-assembly of these species is the inherent information held within the ligand strand. If the ligand is sufficiently instructed or “pre-programmed” it will partition into different binding domains and coordinate two or more metal ions, thus preventing the formation of mononuclear complexes. We have previously shown how ligands that contain a mixture of pyridyl and thiazolyl *N*-donor groups form a variety of helical complexes whose geometries are determined, in part, by the natural break of the ligands into distinct binding domains imposed by the presence of the thiazole units in the oligopyridyl chain.⁷

In this paper, we describe the use of pyridyl *N*-oxides to partition ligands into different binding domains. To the best of our knowledge, this is the first example of the use of *N*-oxide ligands for the formation of transition metal helicates, although the formation of mononuclear complexes from *N*-oxides of 2,2':6',2''-terpyridine and 2,2':6',2'':6'',2'''-quaterpyridine have been reported.^{8,9}

2. Results and discussion

Both ligands **L**¹ and **L**² were synthesized as shown in Scheme 1. The starting material in both cases was 2,2':6',2''-terpyridine tris-*N*-oxide, which was converted to the 6,6''-dicyano derivative by reaction of trimethylsilyl cyanide and benzoyl chloride in refluxing CH_2Cl_2 . Further reaction with hydrogen sulfide gas gave the 6,6''-dithioamide in excellent yield. Reaction of

this compound with either chloroacetone or 2-(α -bromoacetyl)pyridine gave ligands **L**¹ and **L**², respectively.

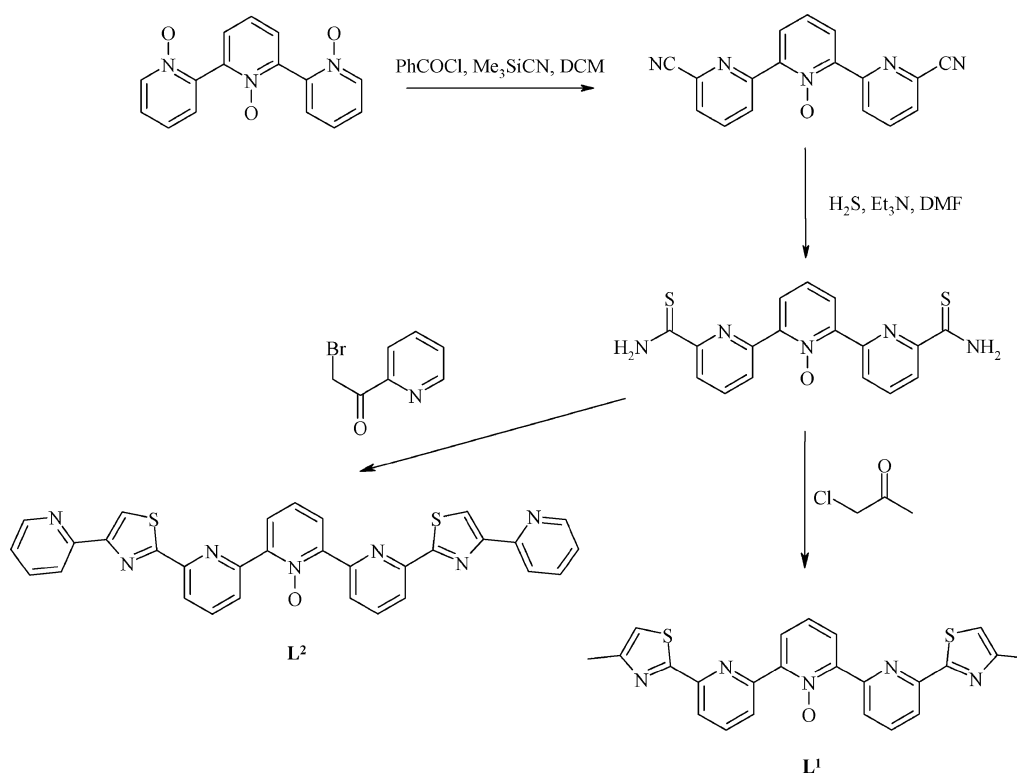
Although the ligand **L**¹ is insoluble in common organic solvents, reaction with $\text{Cd}(\text{ClO}_4)_2 \cdot 6\text{H}_2\text{O}$ in MeCN produced a clear pale yellow solution which gave ions in the ESI-MS at m/z 966 and 1409, corresponding to $\{[\text{Cd}_2(\text{L}^1)](\text{ClO}_4)_3\}^+$ and $\{[\text{Cd}_2(\text{L}^1)_2](\text{ClO}_4)_3\}^+$, indicating the formation of a dinuclear double-stranded helicate. The solid-state structure was confirmed from X-ray crystallographic analysis. In the complex cation, two Cd^{2+} ions are coordinated by two bridging ligands in a double helicate arrangement (Fig. 1 and 2). Both ligands are partitioned into two tridentate binding domains, with the *N*-oxide acting as a bridging bidentate donor to both cadmium(II) ions. The NCCN torsion angles between the terminal pyridyl-thiazole domain and central pyridine *N*-oxide, range from 39(1)–43(1)°. As a consequence of the bridging oxygen atoms in the $\{\text{Cd}(\mu\text{-O}_2)\text{Cd}\}^{4+}$ core, the metal–metal distance (*i.e.* the effective helical pitch length) is quite short (4.027(1) Å) when compared to other cadmium-containing helicates.^{14,15} The $\text{Cd}-\text{O}_{(\text{N-oxide})}$ distances range from 2.327(3)–2.355(3) Å, whilst those to the *N*-donors lie in the range 2.256(4)–2.537(4) Å (Table 1). Each of the two cadmium centres is further coordinated by one oxygen atom of a bridging perchlorate anion. By interacting with the $\text{Cd}(\text{II})$ ion in the neighbouring cation, these interstitial perchlorates link adjacent helicates end-on throughout the lattice. The infinite columns propagate along the crystallographic *c*-axis, but unlike previously observed packing motifs,¹⁶ the helicate units alternate between having *P*- and *M*-screw senses.

In a similar manner to **L**¹, in its non-complexed form **L**² is insoluble in common organic solvents. However, reaction with $\text{Cd}(\text{ClO}_4)_2 \cdot 6\text{H}_2\text{O}$ in MeNO₂ produces a clear pale yellow solution, the ESI-MS of which shows ions at m/z 1662, 1352 and 782, corresponding to $\{[\text{Cd}_2(\text{L}^2)_2](\text{ClO}_4)_3\}^+$, $\{[\text{Cd}(\text{L}^2)_2](\text{ClO}_4)_3\}^+$ and $\{[\text{Cd}(\text{L}^2)](\text{ClO}_4)_3\}^+$ species, respectively. The formation of a dinuclear double helicate $[\text{Cd}_2(\text{L}^2)_2]^{4+}$ in the solid-state was confirmed from single crystal X-ray diffraction studies (Fig. 3). The structure shows two distorted octahedral cadmium cations coordinated by two ligands. Each of the ligands coordinate *via* the central *N*-oxide and the two terminal bidentate pyridyl-thiazole domains. The

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Scheme 1 Synthesis of ligands **L**¹ and **L**².

N-oxide again acts as a bridging bidentate donor coordinating both cadmium cations, with $\text{Cd}-\text{O}$ bond lengths ranging from 2.373(3)–2.395(3) Å (Table 2). The $\text{Cd}-\text{N}$ distances involving the terminal pyridyl-thiazole domains range from 2.299(4) to 2.593(4) Å (Fig. 4). However, the pyridyl units adjacent to the central *N*-oxide of the terpyridyl domain have $\text{Cd}-\text{N}$ distances between 2.732(3)–2.772(3) Å; interactions which are rather too long to be considered bonding. These units may thus be considered as non-bonding or hypodentate. Again the “twist” caused by partitioning of the ligand strands about the central pyridine *N*-oxide evokes NCCN torsion angles in the range 32(1)–38(1)°. The pitch length of this helicate (4.347(1) Å) is slightly longer than that of

$[\text{Cd}_2(\text{L}^1)_2]^{4+}$ but is still quite short when compared to other cadmium-containing helicates.^{14,15}

Comparison of these complexes with others derived from similar ligands and $\text{Cd}(\text{II})$ demonstrate that the *N*-oxide unit is instrumental in partitioning of the ligand. For example, reaction of a ligand strand identical to **L**² (but which does not have a central *N*-oxide unit) results in the formation of a *trinuclear* double helicate.¹⁴ In this case, the ligand partitions into three different binding domains consisting of two terminal pyridyl-thiazole bidentate units and a central terdentate terpyridyl unit. Thus, the three $\text{Cd}(\text{II})$ ions occupy two different coordination environments, with the unique central cadmium ion coordinated by two terpyridyl units (one from each ligand)

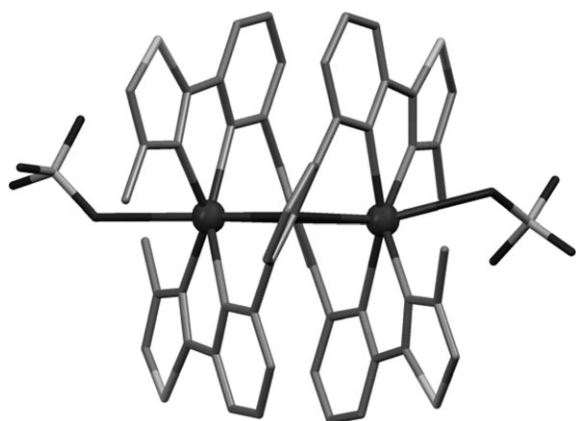


Fig. 1 Structure of the complex cation $[\text{Cd}_2(\text{L}^1)_2(\text{ClO}_4)_2]^{2+}$.

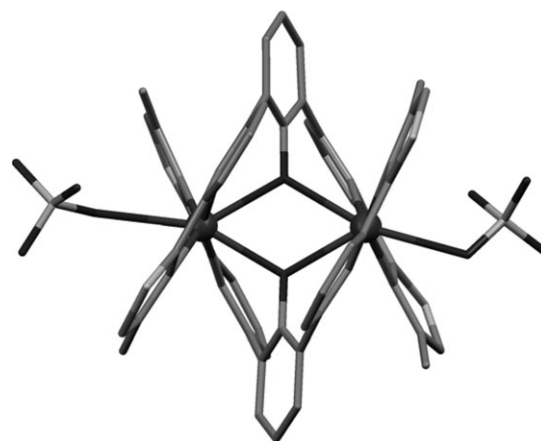


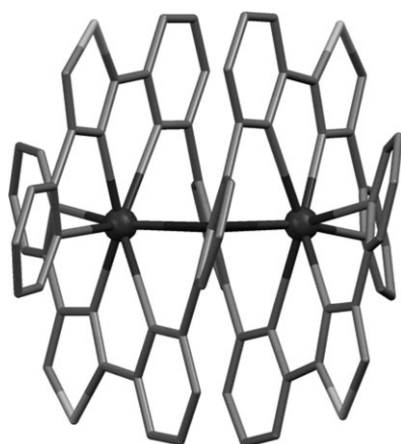
Fig. 2 Structure of the complex cation $[\text{Cd}_2(\text{L}^1)_2(\text{ClO}_4)_2]^{2+}$ showing the $\{\text{Cd}(\mu\text{-O}_2)\text{Cd}\}$ core.

Table 1 Selected bond distances (Å) and angles (°) for $[\text{Cd}_2(\text{L}^1)_2](\text{ClO}_4)_4 \cdot \text{CH}_3\text{CO}_2\text{CH}_2\text{CH}_3$

Cd(1)–N(61)	2.256(4)	Cd(2)–N(51)	2.297(5)
Cd(1)–N(11)	2.277(4)	Cd(2)–N(101)	2.300(4)
Cd(1)–O(81)	2.335(3)	Cd(2)–O(31)	2.327(3)
Cd(1)–O(31)	2.355(3)	Cd(2)–O(81)	2.332(3)
Cd(1)–N(71)	2.452(4)	Cd(2)–N(41)	2.480(4)
Cd(1)–N(21)	2.456(4)	Cd(2)–N(91)	2.537(4)
Cd(1)–O(1B)	2.644(4)	Cd(2)–O(1C)	2.713(4)
N(61)–Cd(1)–N(11)	145.7(1)	N(51)–Cd(2)–N(101)	146.9(2)
N(61)–Cd(1)–O(81)	126.0(1)	N(51)–Cd(2)–O(31)	122.3(2)
N(11)–Cd(1)–O(81)	85.8(1)	N(101)–Cd(2)–O(31)	89.1(1)
N(61)–Cd(1)–O(31)	88.0(1)	N(51)–Cd(2)–O(81)	86.7(1)
N(11)–Cd(1)–O(31)	122.6(1)	N(101)–Cd(2)–O(81)	120.9(1)
O(81)–Cd(1)–O(31)	60.8(1)	O(31)–Cd(2)–O(81)	61.2(1)
N(61)–Cd(1)–N(71)	72.7(1)	N(51)–Cd(2)–N(41)	72.6(2)
N(11)–Cd(1)–N(71)	116.4(1)	N(101)–Cd(2)–N(41)	116.9(1)
O(81)–Cd(1)–N(71)	66.7(1)	O(31)–Cd(2)–N(41)	66.6(1)
O(31)–Cd(1)–N(71)	92.8(1)	O(81)–Cd(2)–N(41)	97.5(1)
N(61)–Cd(1)–N(21)	111.9(1)	N(51)–Cd(2)–N(91)	111.0(1)
N(11)–Cd(1)–N(21)	72.5(1)	N(101)–Cd(2)–N(91)	70.8(1)
O(81)–Cd(1)–N(21)	95.6(1)	O(31)–Cd(2)–N(91)	97.5(1)
O(31)–Cd(1)–N(21)	66.7(1)	O(81)–Cd(2)–N(91)	64.9(1)
N(71)–Cd(1)–N(21)	158.3(1)	N(41)–Cd(2)–N(91)	161.1(1)
N(61)–Cd(1)–O(1B)	76.1(1)	O(1C)–Cd(2)–N(101)	78.0(1)
N(11)–Cd(1)–O(1B)	69.8(1)	O(1C)–Cd(2)–N(51)	73.2(1)
O(81)–Cd(1)–O(1B)	143.0(1)	O(1C)–Cd(2)–N(44)	77.4(1)
O(31)–Cd(1)–O(1B)	156.1(1)	O(1C)–Cd(2)–N(91)	121.4(1)
N(71)–Cd(1)–O(1B)	99.2(1)	O(1C)–Cd(2)–O(31)	130.7(1)
N(21)–Cd(1)–O(1B)	102.5(1)	O(1C)–Cd(2)–O(81)	159.9(1)

and the two identical terminal cadmium ions coordinated by a bidentate pyridyl-thiazole unit. In the case of L^2 , the central terpyridyl unit cannot act as a terdentate unit due to the *N*-oxide, which effectively separates the ligand into two distinct binding domains.

Cadmium complexes of the non-functionalised analogue of L^1 (*i.e.* the same potentially pentadentate ligand but without the *N*-oxide unit) have not been reported to date. However, studies carried out on a similar potentially pentadentate pyridyl-thiazole ligand, containing a pyridyl-thiazole-pyridyl-thiazole-pyridyl chain, showed the formation of a mononuclear complex in which the cadmium ion is coordinated by a near-planar pentadentate “equatorial” ligand.¹⁵ Although it cannot be assumed that the non-functionalised analogue of L^1 would act in a similar fashion, it may be reasonable to expect that, as

**Fig. 3** Structure of the complex cation $[\text{Cd}_2(\text{L}^2)_2]^{4+}$.**Table 2** Selected bond distances (Å) and angles (°) for $[\text{Cd}_2(\text{L}^2)_2](\text{ClO}_4)_4 \cdot 3\text{CH}_3\text{NO}_2$

Cd(1)–N(91)	2.299(4)	Cd(2)–N(61)	2.326(3)
Cd(1)–N(21)	2.315(4)	Cd(2)–N(131)	2.331(4)
Cd(1)–O(111)	2.373(3)	Cd(2)–O(41)	2.388(3)
Cd(1)–O(41)	2.377(3)	Cd(2)–O(111)	2.395(3)
Cd(1)–N(11)	2.558(4)	Cd(2)–N(141)	2.478(4)
Cd(1)–N(81)	2.593(4)	Cd(2)–N(71)	2.580(4)
N(91)–Cd(1)–N(21)	153.1(2)	N(61)–Cd(2)–N(131)	153.2(1)
N(91)–Cd(1)–O(111)	116.4(1)	N(61)–Cd(2)–O(41)	117.1(1)
N(21)–Cd(1)–O(111)	88.4(1)	N(131)–Cd(2)–O(41)	87.6(1)
N(91)–Cd(1)–O(41)	87.4(1)	N(61)–Cd(2)–O(111)	88.1(1)
N(21)–Cd(1)–O(41)	114.7(1)	N(131)–Cd(2)–O(111)	114.1(1)
O(111)–Cd(1)–O(41)	62.6(1)	O(41)–Cd(2)–O(111)	62.1(1)
N(91)–Cd(1)–N(11)	99.6(1)	N(61)–Cd(2)–N(141)	98.5(1)
N(21)–Cd(1)–N(11)	67.9(1)	N(131)–Cd(2)–N(141)	68.4(1)
O(111)–Cd(1)–N(11)	91.3(1)	O(41)–Cd(2)–N(141)	92.6(1)
O(41)–Cd(1)–N(11)	153.1(1)	O(111)–Cd(2)–N(141)	153.9(1)
N(91)–Cd(1)–N(81)	67.4(1)	N(61)–Cd(2)–N(71)	66.8(1)
N(21)–Cd(1)–N(81)	93.3(1)	N(131)–Cd(2)–N(71)	95.2(1)
O(111)–Cd(1)–N(81)	159.0(1)	O(41)–Cd(2)–N(71)	156.0(1)
O(41)–Cd(1)–N(81)	98.0(1)	O(111)–Cd(2)–N(71)	95.3(1)
N(11)–Cd(1)–N(81)	108.7(1)	N(141)–Cd(2)–N(71)	110.6(1)

with L^2 , the *N*-oxide ring of the central terpyridyl unit will prevent the ligand from efficiently coordinating to one metal ion.

3. Experimental

3.1 General details

All reagents and starting materials were purchased from Aldrich and/or Lancaster and were used without further purification. ^1H NMR spectra were obtained on a Bruker 400 MHz NMR instrument and mass spectrometry was performed on a Micromass Quattro VG quadrupole and a Bruker Micro-TOF LC mass spectrometer. 2,2',6',2''-terpyridine-1,1',1''-trioxide was prepared *via* oxidation with H_2O_2 in acetic acid.¹⁰

3.2 Synthesis

3.2.1 Synthesis of 2,2',6',2''-terpyridine-6,6''-dicarbonitrile-1'-oxide. To a suspension of 2,2',6',2''-terpyridine-1,1',1''-trioxide (0.4 g, 1.4 mmol) in dichloromethane (30 cm^3), benzoyl

**Fig. 4** Structure of the complex cation $[\text{Cd}_2(\text{L}^2)_2]^{4+}$ showing the $\{\text{Cd}(\mu\text{-O}_2)\text{Cd}\}$ core.

chloride (0.66 g, 4.7 mmol) and trimethylsilyl cyanide (0.78 g, 7.9 mmol) were added and the solution refluxed for 20 h. The resulting white solid was isolated *via* filtration giving the 6,6''-dicyano derivative. Yield: 0.41 g, 96%. ^1H NMR [400 MHz, $(\text{CD}_3)_2\text{SO}$, 70 °C]: δ (ppm) 8.79 (2H, d, J = 7.47; py), 8.19 (2H, dd, J = 8.00, 8.00; py), 8.09 (4H, d, J = 7.90; py), 7.66 (1H, t, J = 7.96 Hz; py). ESI-MS: m/z 300.1 $[\text{M} + \text{H}^+]$.

3.2.2 Synthesis of 6,6''-dithioamide-2,2',6',2''-terpyridine-1'-oxide. 2,2',6',2''-Terpyridine-6,6''-dicarbonitrile-1'-oxide (0.41 g, 1.4 mmol) was dissolved in a mixture of DMF (20 cm^3) and triethylamine (2 cm^3), and H_2S gas was slowly bubbled through this solution and a brown/beige precipitate was formed. The solution was filtered and concentrated. Addition of ethanol (5 cm^3) gave a yellow precipitate which was isolated by filtration giving 6,6''-dithioamide-2,2',6',2''-terpyridine-1'-oxide as a yellow solid. Yield: 0.35 g, 70%. ^1H NMR [400 MHz, $(\text{CD}_3)_2\text{SO}$, 70 °C]: δ (ppm) 10.31 (2H, s; $-\text{NH}_2$), 10.05 (2H, s; $-\text{NH}_2$), 8.83 (2H, d, J = 8.00; py), 8.56 (2H, d, J = 7.89; py), 8.47 (2H, d, J = 7.97; py), 8.11 (2H, dd, J = 7.99, 7.99; py), 7.62 (1H, t, J = 8.00 Hz; py). ESI-MS: m/z 390.1 $[\text{M} + \text{Na}^+]$.

3.2.3 Synthesis of L^1 . The 6,6''-dithioamide-2,2',6',2''-terpyridine-1'-oxide (0.1 g, 0.27 mmol) and chloroacetone (0.05 g, 0.54 mmol) were heated to 80 °C in anhydrous DMF (25 cm^3) for 2 h and a brown precipitate was formed. The precipitate was collected by filtration and washed with ether (2 cm^3) to give the ligand L^1 as its hydrochloride salt. The free base was prepared by suspending the $\text{L}^1 \cdot \text{HCl}$ in concentrated ammonia (0.88 sg, 20 cm^3) and leaving the mixture for 12 h. The suspension was filtered and the solid was washed with ethanol (2 cm^3) and diethyl ether (2 cm^3) to give the ligand L^1 as a beige solid. Yield: 0.073 g, 62%. $\text{L}^1 \cdot \text{HCl}$: ^1H NMR [400 MHz, $(\text{CD}_3)_2\text{SO}$, 70 °C]: δ (ppm) 8.68 (2H, d, J = 7.72; py), 8.18 (4H, d, J = 7.30; py), 8.09 (2H, dd, J = 7.89, 7.89; py), 7.70 (1H, t, J = 7.90 Hz; py), 7.42 (2H, s; th), 2.51 (6H, s; $-\text{CH}_3$). ESI-MS: m/z 443.1 $[\text{M}^+]$.

3.2.4 Synthesis of L^2 . The 6,6''-dithioamide-2,2',6',2''-terpyridine-1'-oxide (0.050 g, 0.14 mmol) and 2-(α -bromoacetyl)-pyridine (0.051 g, 0.25 mmol) were heated to 80 °C in DMF (20 cm^3) for 2 h and a brown precipitate was formed. The precipitate was collected by filtration and washed with ethanol (3 cm^3) and diethyl ether (3 cm^3) to give the ligand L^2 as its hydrobromide salt. The free base was prepared by suspending the $\text{L}^2 \cdot \text{HBr}$ in concentrated ammonia (0.88 sg, 20 cm^3) and leaving the mixture for 12 h. The suspension was filtered and the solid was washed with ethanol (2 cm^3) and diethyl ether (2 cm^3) to give the ligand L^2 as a brown solid. Yield: 0.039 g, 49%. $\text{L}^2 \cdot \text{HBr}$: ^1H NMR [400 MHz, $(\text{CD}_3)_2\text{SO}$, 70 °C]: δ (ppm) 8.75 (2H, py), 8.69 (2H, py), 8.48 (2H, s; th), 8.37 (2H, py), 8.28 (2H, py), 8.24 (2H, py), 8.18 (2H, py), 8.02 (2H, py), 7.75 (1H, py), 7.45 (2H, py). ESI-MS: m/z 569.1 $[\text{M}^+]$.

3.2.5 Preparation of $[\text{Cd}_2(\text{L}^1)_2](\text{ClO}_4)_4$. To a suspension of L^1 (0.005 g, 0.011 mmol) in acetonitrile (2 cm^3), $\text{Cd}(\text{ClO}_4)_2 \cdot 6\text{H}_2\text{O}$ (0.0046 mg, 0.011 mmol) was added and the solution was stirred until dissolution was complete. Filtration followed by slow diffusion of diethyl ether into the

solution gave $[\text{Cd}_2(\text{L}^1)_2](\text{ClO}_4)_4$ as colourless crystals. Yield: 0.0079 g, 47%. ESI-MS: m/z 1409 $\{[\text{Cd}_2(\text{L}^1)_2](\text{ClO}_4)_3\}^+$. Elemental analysis calculated for $\text{C}_{46}\text{H}_{34}\text{N}_{10}\text{Cd}_2\text{Cl}_4\text{O}_{18}\text{S}_4$: C, 36.6; H, 2.3; N, 9.3%. Found: C, 36.4; H, 2.5; N, 8.8%.

3.2.6 Preparation of $[\text{Cd}_2(\text{L}^2)_2](\text{ClO}_4)_4$. To a suspension of L^2 (0.010 g, 0.018 mmol) in nitromethane (2 cm^3), $\text{Cd}(\text{ClO}_4)_2 \cdot 6\text{H}_2\text{O}$ (0.0075 g, 0.018 mmol) was added and the solution was stirred until dissolution was complete. Filtration followed by slow diffusion of diethyl ether into the solution gave $[\text{Cd}_2(\text{L}^2)_2](\text{ClO}_4)_4$ as colourless crystals. Yield: 0.008 g, 51%. ESI-MS: m/z 1662 $\{[\text{Cd}_2(\text{L}^2)_2](\text{ClO}_4)_3\}^+$. Elemental analysis calculated for $\text{C}_{62}\text{H}_{38}\text{N}_{14}\text{Cd}_2\text{Cl}_4\text{O}_{18}\text{S}_4 \cdot \text{MeNO}_2$: C, 41.5; H, 2.3; N, 11.5%. Found: C, 41.9; H, 2.3; N, 11.1%.

3.3 Crystallography†

Diffraction intensity data were collected on a Bruker SMART CCD area-detector diffractometer with Mo-K α radiation (λ = 0.71073 Å). Intensities were integrated¹¹ from several series of exposures, each exposure covering 0.3° in ω , and the total data set being a sphere. Absorption corrections were applied based on multiple and symmetry-equivalent measurements.¹² The structure was solved by direct methods and refined by least squares on weighted F^2 values for all reflections.¹³

3.3.1 Crystal data for $[\text{Cd}_2(\text{L}^1)_2](\text{ClO}_4)_4 \cdot \text{CH}_3\text{CO}_2\text{CH}_2\text{CH}_3$. $\text{C}_{50}\text{H}_{42}\text{Cd}_2\text{Cl}_4\text{N}_{10}\text{O}_{20}\text{S}_4$, M = 1597.78, monoclinic, space group $P2(1)/c$, a = 23.568(4), b = 12.894(2), c = 21.345(4) Å, β = 110.065(4)°, V = 6093(2) Å³, Z = 4, D_{calc} = 1.742 Mg m⁻³, $F(000)$ = 3200, μ = 1.093 mm⁻¹, crystal dimensions 0.4 × 0.3 × 0.2 mm, T = 173(2) K. A total of 64 281 reflections were measured in the range $1.83 \leq \theta \leq 27.52^\circ$ (hkl range indices $-30 \leq h \leq 30$, $-16 \leq k \leq 16$, $-27 \leq l \leq 27$), 13 984 independent reflections (R_{int} = 0.0539). The structure was refined to F^2 to $wR2$ = 0.1615, $R1$ = 0.0548 [8783 reflections with $I > 2\sigma(I)$] and GOF = 0.992 on F^2 for 812 refined parameters, largest difference peak and hole 1.603 and -1.600 eÅ^{-3} .

3.3.2 Crystal data for $[\text{Cd}_2(\text{L}^2)_2](\text{ClO}_4)_4 \cdot 3\text{CH}_3\text{NO}_2$. $\text{C}_{65}\text{H}_{47}\text{Cd}_2\text{Cl}_4\text{N}_{17}\text{O}_{24}\text{S}_4$, M = 1945.04, triclinic, space group $P1$, a = 11.4892(8), b = 13.6419(9), c = 23.7577(16) Å, α = 93.2420(10), β = 99.9830(10), γ = 97.9550(10)°, V = 3619.6(4) Å³, Z = 2, D_{calc} = 1.785 Mg m⁻³, $F(000)$ = 1952, μ = 0.944 mm⁻¹, crystal dimensions 0.4 × 0.2 × 0.2 mm, T = 173(2) K. A total of 27 891 reflections were measured in the range $1.51 \leq \theta \leq 27.62^\circ$ (hkl range indices $-14 \leq h \leq 14$, $-17 \leq k \leq 14$, $-30 \leq l \leq 30$), 15 653 independent reflections (R_{int} = 0.0411). The structure was refined to F^2 to $wR2$ = 0.0927, $R1$ = 0.0439 [9053 reflections with $I > 2\sigma(I)$] and GOF = 0.906 on F^2 for 1111 refined parameters, largest difference peak and hole 1.902 and -1.011 eÅ^{-3} .

4. Conclusions

The syntheses of two symmetric polydentate pyridyl-thiazolyl ligands, whose central unit comprises a pyridine *N*-oxide

† CCDC 654463, $[\text{Cd}_2(\text{L}^1)_2]^{4+}$, and 654464, $[\text{Cd}_2(\text{L}^2)_2]^{4+}$. For crystallographic data in CIF or other electronic format, see DOI: 10.1039/b711039a

moiety, have been described. Reaction of either of these ligands with $\text{Cd}(\text{ClO}_4)_2 \cdot 6\text{H}_2\text{O}$ results in the formation of a dinuclear double-stranded helicate complex. Comparison of these complexes with those formed from analogous and/or related non-functionalised ligands, demonstrates how *N*-oxide functionalisation of the central pyridine ring induces the ligand to partition into two discrete binding domains.

Acknowledgements

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